

Write the formula or the name for the following compounds.

1. copper(II) bromate

13. PbO_2

2. dichlorine heptoxide

14. S_2O_5

3. copper(II) sulfide

15. PbO

4. dihydrogen dioxide

16. N_2O

5. copper(II) phosphate

17. $(\text{NH}_4)_3\text{PO}_3$

6. nitrogen monoxide

18. CCl_4

7. lead(IV) sulfite

19. $(\text{NH}_4)_2\text{S}$

8. dihydrogen sulfide

20. ClO

9. lead(IV) acetate

21. $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$

10. hydrogen chloride

22. H_2O

11. lead(IV) phosphide

23. BaO

12. sulfur trioxide

24. PF_3

Remember that for molecular compounds, no charges are involved, use the prefixes to determine the number of each atom. For ionic, do not use prefixes unless they are a part of the name of the polyatomic ion.

1. ionic - $\text{Cu}(\text{BrO}_3)_2$
2. molecular - Cl_2O_7
3. ionic - CuS
4. molecular - H_2O_2
5. ionic - $\text{Cu}_3(\text{PO}_4)_2$
6. molecular - NO
7. ionic - $\text{Pb}(\text{SO}_3)_2$
8. molecular - H_2S
9. ionic - $\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_4$
10. molecular - HCl (also called hydrochloric acid)
11. ionic - Pb_3P_4
12. molecular - SO_3
13. ionic - lead(IV) oxide
14. molecular - disulfur pentoxide
15. ionic - lead(II) oxide
16. molecular - dinitrogen oxide
17. ionic - ammonium phosphite
18. molecular - carbon tetrachloride
19. ionic - ammonium sulfide
20. molecular - chlorine oxide (also called chlorous oxide)
21. ionic - ammonium acetate
22. molecular - dihydrogen oxide ... water!
23. ionic - barium oxide
24. molecular - phosphorous trifluoride