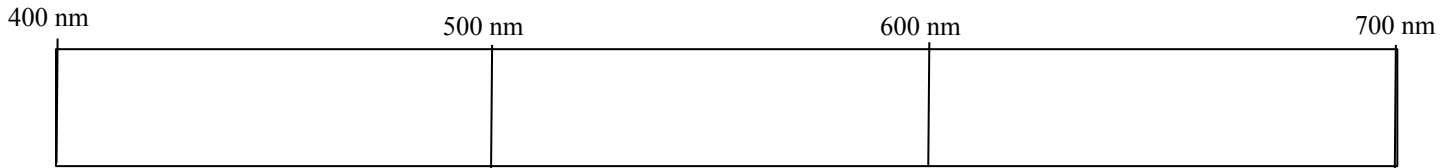
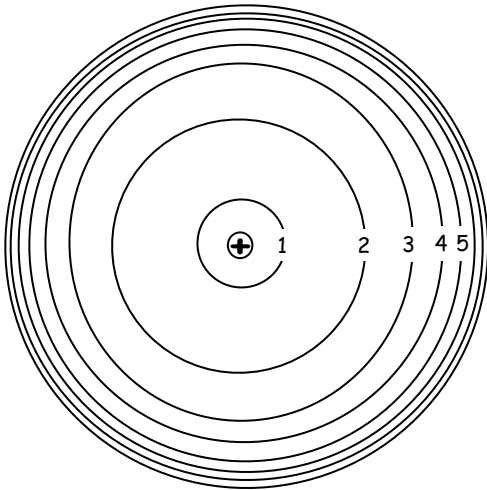


1. Make a sketch of the atomic bright-line emission spectrum for hydrogen.



2. In the atomic model below, represent the cause of each line in the hydrogen spectrum that you drew above. Use a color pencil to represent the appropriate colors.



7. What is it that makes the emission spectra all so different (and unique) for different elements?

After looking at the emission spectra of helium and other elements through the spectroscope, draw sketch of the atomic bright-line emission spectrum for these elements. Use the numbers inside the spectroscope mark the approximate wavelength of each line on the diagram below.

