

P.A.1 (pg 1 of 1) **Nomenclature: Ionic and Molecular Compounds & Acids**

Name _____ Per _____

- Write the formula or the name for the following compounds. Try to work just with a periodic table
- Fold the page over, or cover it with a blank sheet to avoid "cheating".
- The more you practice, the sooner you'll learn the polyatomic ions.

1. copper(II) bromate
2. dichlorine heptoxide
3. cobalt(II) sulfide
4. dihydrogen dioxide
5. nitrogen monoxide
6. silver nitrite
7. lead(II) sulfite
8. hydrosulfuric acid
9. lead(IV) acetate
10. hydrochloric acid
11. chloric acid
12. chlorous acid
13. tin(IV) phosphide
14. sulfur trioxide
15. manganese(II) phosphate
16. MgO
17. HCN
18. PbO₂
19. S₂O₅
20. KMnO₄
21. N₂O
22. (NH₄)₃PO₃
23. CCl₄
24. Zn(OH)₂
25. (NH₄)₂S
26. ClO
27. HBrO₃
28. HBr
29. HBrO₂
30. HBrO
31. NH₄C₂H₃O₂
32. H₂O
33. PF₃
34. HIO₃
35. Cu₂O
36. Cr(CO₃)₃

ANSWERS

- Remember that for molecular compounds, no charges are involved, use the prefixes to identify the number of each atom.
- For ionic, do not use prefixes unless they are a part of the name of the polyatomic ion.

1. ionic – Cu(BrO₃)₂
2. molecular – Cl₂O₇
3. ionic – CoS
4. molecular – H₂O₂
5. molecular – NO
6. ionic – AgNO₂
7. ionic – PbSO₃
8. acid – H₂S
9. ionic – Pb(C₂H₃O₂)₄
10. acid – HCl
11. acid – HClO₃
12. acid – HClO₂
13. ionic – Sn₃P₄
14. molecular – SO₃
15. ionic – Mn₃(PO₄)₂
16. ionic – magnesium oxide
17. hydrocyanic acid
18. ionic – lead(IV) oxide
19. molecular – disulfur pentoxide
20. ionic – potassium permanganate
21. molecular - dinitrogen oxide
22. ionic – ammonium phosphite
23. molecular – carbon tetrachloride
24. ionic – zinc hydroxide
25. ionic – ammonium sulfide
26. molecular – chlorine oxide or chlorine monoxide (also called chlorous oxide)
27. bromic acid
28. hydrobromic acid
29. bromous acid
30. hypobromous acid
31. ionic – ammonium acetate
32. molecular – dihydrogen oxide or dihydrogen monoxide... aka water!
33. molecular – phosphorous trifluoride
34. iodic acid
35. copper(I) oxide
36. chromium(VI) carbonate